

Acid Nomenclature

Acids are substances which produce H^+ ions in water. We name the acids based on the negative ion which is left over after the H^+ ion is given off.

Negative ion suffix

___ate
___ite
___ide

Acid

___ic acid
___ous acid
hydro ___ic acid

Learn the names and formulas for these common acids:

| Formulas | | Names |
|-------------------|--------------------|-------------------|
| H_2SO_4 (aq) | (sulfate ion) | Sulfuric acid |
| H_2SO_3 (aq) | (sulfite ion) | Sulfurous acid |
| HNO_3 (aq) | (nitrate ion) | Nitric acid |
| HNO_2 (aq) | (nitrite ion) | Nitrous acid |
| H_3PO_4 (aq) | (phosphate ion) | Phosphoric acid |
| $HClO_4$ (aq) | (perchlorate ion) | Perchloric acid |
| $HClO_3$ (aq) | (chlorate ion) | Chloric acid |
| $HClO_2$ (aq) | (chlorite ion) | Chlorous acid |
| $HClO$ (aq) | (hypochlorite ion) | Hypochlorous acid |
| HCl (aq) | (chloride ion) | Hydrochloric acid |
| HF (aq) | (fluoride ion) | Hydrofluoric acid |
| HBr (aq) | (bromide ion) | Hydrobromic acid |
| | | |
| HI (aq) | (iodide ion) | Hydroiodic acid |
| $HC_2H_3O_2$ (aq) | (acetate ion) | Acetic Acid |
| HCN (aq) | (cyanide ion) | Hydrocyanic acid |
| H_2CO_3 (aq) | (carbonate ion) | Carbonic Acid |